

31. The property which increases upto group IV-A then decreases onwards

- [A.](#) Ionization energy
- [B.](#) Atomic radii
- [C.](#) Melting & boiling points
- [D.](#) Atomic volume

Answer: Option c

32. Which of the following ion is stable in aqueous solution?

- [A.](#) H⁺
- [B.](#) H⁻
- [C.](#) Cl⁻
- [D.](#) All are stable

Answer: Option c

33. The atoms of same element having same atomic number but different mass number are called

- [A.](#) Isobars
- [B.](#) Isomers
- [C.](#) Isotopes
- [D.](#) Isotropes

Answer: Option c

34. Deuterium reacts with oxygen to form

- [A.](#) Hard water
- [B.](#) Heavy water
- [C.](#) Soft water
- [D.](#) Water gas

Answer: Option B

35. Which order of ionization energy is correct

- [A.](#) Mg < Al
- [B.](#) Si > P
- [C.](#) Mg > Al
- [D.](#) both b & c

Answer: Option c

36. Ionization energy depends upon

- A. Nuclear charge
- B. Atomic size
- C. Shielding effect
- D. I.E depends upon all of the above and nature of orbital

Answer: Option D

37. Shielding effect across the period

- A. Increases
- B. Decreases
- C. Can not be predicted
- D. Remains constant

Answer: Option D

38. Addition of 2nd electron to a uninegative ion is always

- A. Exothermic
- B. Endothermic
- C. Data is insufficient
- D. Unpredictable

Answer: Option B

39. Higher value of electron affinity means

- A. Atom will lose electron easily
- B. Atom will gain electron easily
- C. Atom may form di-positive ion
- D. The reason is unknown

Answer: Option B

40. Metallic characters of alkali metals

- A. Increase down the group
- B. Decrease down the group
- C. No regular trend
- D. Remain same

Answer: Option A