

21. Amphoteric oxides are those which possess \_\_\_\_\_ properties

- A. Acidic B. Basic  
C. Acidic and basic D. Neutral and acidic

**Answer:** Option C

22. Best position of hydrogen in the periodic table is above I.A Group which is mainly due to

- A. Both are electropositive B. Similar outer most shell electronic configuration  
C. Both form ionic compounds D. All

**Answer:** Option D

23. Hydrogen resembles with carbon because of having

- A. Same number of electrons in the valence shell B. Similar physical state  
C. Remarkable reducing properties D. Homovalent (show same valency)

**Answer:** Option C

24. Which one of the following sets consists of all coinage metals?

- A. Cu Hg Au B. Cu Ag Au  
C. Ag Au Hg D. Cu Fe Au

**Answer:** Option B

25. In which of the following pairs are elements belonging to the same group?

- A. Boron & Beryllium B. Nitrogen & Phosphorous  
C. Magnesium & Aluminium D. Gallium & Helium

**Answer:** Option B

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26. Many properties of an element and its compounds can be predicted from the position of the element in the periodic table. What property could not be predicted in this way?

- A. The nature of its oxides                      B. The charge on its ions  
C. The formula of its oxide                      D. Its number of isotopes

**Answer:** Option D

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27. Which one of the following is not a periodic property?

- A. Melting point of elements                      B. Boiling point of elements  
C. Ionization energy of elements                      D. Coordination number of ions

**Answer:** Option D

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28. The atomic radii decreases by increasing atomic number in

- A. Alkali metal                      B. Alkaline earth metal  
C. Elements from Li to Ne                      D. Halogens

**Answer:** Option C

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29. Which discovery caused a revision in the periodic law as stated by Mendeleev?

- A. Location of nucleus by Rutherford                      B. Atomic number by Moseley  
C. X-rays by Roentgen                      D. Natural radioactivity by Henry Bacquerel.

**Answer:** Option B

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30. An element has electronic configuration  $1s^2 2s^2 2p^2$ . It belongs to

- A. Group II-A                      B. Group IV-A  
C. Group VII-A                      D. Group VI-A

**Answer:** Option B

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