

21. Mixtures which distill over without change in composition called
- A. zeotropic mixture B. azeotropic mixture
C. amphoteric mixture D. ideal solution
- Answer:** Option B
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22. Concentration of solute molecule when they are in equilibrium with solid substance at particular temperature is called
- A. saturated solution B. solubility
C. unsaturated solution D. super saturated solution
- Answer:** Option B
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23. Solubility of $KClO_3$ gives
- A. continuous and falling solubility curve B. discontinuous and falling solubility curve
C. continuous and rising solubility curve D. discontinuous and rising solubility curve
- Answer:** Option C
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24. The determination of correct molecular weight from Raoult's law is applicable to
- A. a volatile solute in dilute solution B. a non-electrolyte & non volatile solute in concentrated solution
C. a non-electrolyte & non volatile solute in concentrated solution D. non volatile solute in a dilute solution
- Answer:** Option D
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25. Boiling point elevations can be measured by
- A. Beckmann's method B. Landsbergers method
C. Linds method D. none of the above
- Answer:** Option B

26. Beckmanns apparatus is used to measure

- A. boiling point elevation
- B. depression in freezing point
- C. lowering of vapour pressure
- D. lowering of osmotic pressure

Answer: Option **B**

27. Water molecules surrounds more around

- A. ve ion
- B. complex ion
- C. ?ve ion
- D. neutral atom

Answer: Option **A**

28. The compounds in which water molecules are added are called

- A. Hydrated ions
- B. double salts
- C. hydrates
- D. complexes

Answer: Option **C**

29. Hydration is a process in which

- A. Molecules are surrounded by solvent molecules
- B. Ions are surrounded by solvent molecules
- C. Both ions and molecules are surrounded by solvent molecules
- D. Both ions and molecules are surrounded by water molecules

Answer: Option **D**

30. Solution of Na₂SO₄ will be

- A. basic
- B. acidic
- C. neutral
- D. cannot be predicted without data

Answer: Option **C**