

1. The rate of reaction

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|---|---|
| <u>A.</u> Increases as the reaction proceeds | <u>B.</u> Decreases as the reaction proceeds |
| <u>C.</u> Remains the same as the reaction proceeds | <u>D.</u> May decrease or increase as the reaction proceeds |

Answer: Option B

2. The addition of a catalyst to the reaction system

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| <u>A.</u> Increases the rate of forward reaction only | <u>B.</u> Increases the rate of reverse reaction |
| <u>C.</u> Increases the rate of forward but decreases the rate of backward reaction | <u>D.</u> Increases the rate of forward as well as backward reaction equally |

Answer: Option D

3. The specific rate constant of a first order reaction depends on the

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| <u>A.</u> Time | <u>B.</u> Concentration of the reactant |
| <u>C.</u> Temperature | <u>D.</u> Concentration of the product |

Answer: Option C

4. On increasing the temperature the rate of reaction increases mainly because

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|---|---|
| <u>A.</u> The activation energy of the reaction increases | <u>B.</u> Concentration of the reacting molecules increases |
| <u>C.</u> Collision frequency increases | <u>D.</u> None of these |

Answer: Option C

5. $\frac{d[N_2O_5]}{dt}$ represents

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|---|---|
| <u>A.</u> Rate of formation of N_2O_5 | <u>B.</u> Rate of decomposition of N_2O_5 |
| <u>C.</u> order of the reaction | <u>D.</u> none |

Answer: Option B

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6. The value of activation energy is primarily determined by
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|--------------------------------------|---|
| <u>A.</u> Temperature | <u>B.</u> Effective collision |
| <u>C.</u> Concentration of reactants | <u>D.</u> Chemical nature of reactants and products |

Answer: Option B

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7. Sum of exponents of molar concentration is called
- | | |
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| <u>A.</u> Order of reaction | <u>B.</u> Molecularity |
| <u>C.</u> Rate of reaction | <u>D.</u> Average of reaction |

Answer: Option A

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8. Spontaneous reactions are
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| <u>A.</u> Moderate | <u>B.</u> Slow |
| <u>C.</u> Fast | <u>D.</u> not natural |

Answer: Option C

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9. In rate expression the concentration of reactants is negative. It shows
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|---|---|
| <u>A.</u> Concentration of reactant does not change | <u>B.</u> Concentration of product increases |
| <u>C.</u> Concentration of reactant decreases | <u>D.</u> Concentration of reactant increases |

Answer: Option C

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10. Unit of rate of reaction is
- | | |
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| <u>A.</u> Moles $\text{dm}^{-3} \text{sec}^{-1}$ | <u>B.</u> Moles dm^{-3} |
| <u>C.</u> Moles sec^{-1} | <u>D.</u> $\text{Mol}^{-1} \text{dm}^3 \text{sec}^{-1}$ |

Answer: Option A

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