

1. The relative attraction of the nucleus for the electrons in a chemical bond is called
- A. Ionization energy B. Electron affinity
C. Electro negativity D. None of the above

Answer: Option C

2. The ionization energy
- A. Generally increases from left to right in a period B. Does not change in a period
C. Increase from top to bottom in a group D. Does not change in a group

Answer: Option A

3. Which of the following will have highest value of electron affinity
- A. F B. Cl
C. Br D. I

Answer: Option B

4. Which type of bond is formed by overlap of p orbitals
- A. Pi (?) B. Sigma(?)
C. Both D. Neither

Answer: Option C

5. The octet rule does not always hold for which of the following elements
- A. C B. O
C. F D. P

Answer: Option D

